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solvent

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 explanation the molarity of a mixture  
 mmix can be calculated using the  
 following formula  $m_{mix} = \frac{m_1 v_1 + m_2 v_2}{v_1 + v_2}$   
 $\frac{1.5 \times 480 + 1.2 \times 520}{480 + 520}$   
 $\frac{720 + 624}{1000} = 1.344$  m q5 with increase  
 in temperature which of these changes  
 a molality b fraction of solute present in  
 water

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 solution which contains 0.40 mol of a  
 substance dissolved in 1.6 l of a  
 solution 2 what is the molarity of a  
 solution containing 325 g of nacl  
 dissolved in 750 ml of solution 3 140 g  
 of kcl is dissolved in 600 ml of water  
 what is the molarity 4 724 g of  
 ammonium phosphate in 4500 ml of  
 alcohol

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 the following solutions 0.444 mol of  
 cocl<sub>2</sub> in 0.654 l of solution 98.0 g of

phosphoric acid h<sub>3</sub>po<sub>4</sub> in 1.00 l of  
 solution 0.2074 g of calcium hydroxide  
 ca(oh)<sub>2</sub> in 40.00 ml of solution 10.5 kg  
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