

Narcissistic Mothers And Grown Up Daughters

Christian G. Meyer

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found to be 0.00352 moles of butane
in order to find the molarity the mass of
the butane should be divided by the
moles after further calculations the
molarity of the butane gas was
calculated to be 56.81 grams moles this
lab is done with tons of sources of error
lab report mass of butane pdf gases

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web this data shows that there was a
gain the mass of the lighter which
shows that there was a certain amount
of butane 0.38 grams 0.02 grams gas
that was successfully collected during
the gas collection there is a mild
influence of barometric pressure during
both trials as shown above in the raw
data

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web 11 use avogadro's law to
determine the number of moles of
butane gas assume that butane is an
ideal gas and that one mole has a
volume of 22.4 l at stp $V_1 n_1 = V_2 n_2$
0.896 l n 22.4 l 1 mole n 0.0040 moles
12 experimental molar mass g mole of
butane 0.23 g 0.0040 moles 58 g mole
13 the molecular formula of butane is C_4H_{10}

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web nov 3 2020 in this experiment the
moles and mass will be required to
determine the molar mass of butane in
a butane lighter experimental to begin
the experiment the mass of a butane
lighter was measured before using the

butane gas to deplete the water to the
80ml mark on a graduated cylinder

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web dec 18 2020 in this lab students
will experimentally determine the
molar mass of a gas specifically butane
 C_4H_{10} by collection over water this
experiment is an inquiry based
experiment for 2nd year chemistry or
ap chemistry students who have
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 web the pressure of the gas the pressure of the gas was calculated

using dalton s law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760mm hg the pressure of the water vapor was found by using the temperature of the water
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 number of moles of butane gas assume
 that butane is an ideal gas and that one
 mole has a volume of 22 l at stp
 experimental molar mass g mole of
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 is c 4 h 10

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web although the most convenient gas
 for this is butane other gases may be
 available in similar small cylinders the
 collection of a gas sample and the
 weighing of the gas cylinder before and
 after this should take about 5 10
 minutes as a demonstration

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