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web using algebra the moles was founded to be 0 00352 moles of butane in order to find the molarity the mass of the butane should be divided by the moles after further calculations the molarity of the butane gas was calculated to be 56 81 grams moles this lab is done with tons of sources of error
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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a

volume of 22.4 l at stp v 1 n 1 v 2 n 2 0
0896 l n 22.4 l 1 mole n 0.0040 moles
12 experimental molar mass g mole of
butane 0.23 g 0.0040 moles 58 g mole
13 the molecular formula of butane is C
4 H 10

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web nov 3 2020 in this experiment the
moles and mass will be required to
determine the molar mass of butane in
a butane lighter experimental to begin
the experiment the mass of a butane
lighter was measured before using the
butane gas to deplete the water to the
80ml mark on a graduated cylinder

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web dec 18 2020 in this lab students
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C 4 H 10 by collection over water this
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will not dissolve in the water it is
possible for butane gas possible to form
under water and get trapped possible
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web the pressure of the gas the
pressure of the gas was calculated
using dalton s law of partial pressures
by subtracting the pressure of the
water vapor from the total pressure
which in this case was the atmospheric
pressure 760mm hg the pressure of the
water vapor was found by using the
temperature of the water

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and t the particles in an ideal gas do
not have finite size and volume the
collisions between the ideal gas
particles are said to be elastic they
exert no attractive or repulsive forces

hydrogen gas generated in today s experiment is however a real gas not an ideal gas

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web in this lab you will collect a measurable volume of butane gas using water displacement using dalton s law and the ideal gas law the number of

moles of butane gas and then the molar mass of butane can be calculated your experimental value for molar mass will be compared to the known molar mass of butane and percent error calculated **butane gas chemistry lab report answers secure4 khronos** - Dec 28 2021

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and

after this should take about 5 10 minutes as a demonstration

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