

Meaning Of Isaiah 6 1 8

Xiaolong Qi

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 $2 \text{ CO}_2(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ Δh 94.5 kJ
 $\text{H}_2\text{O}(\text{g}) \rightarrow \frac{1}{2} \text{O}_2(\text{g})$ Δh 71.2 kJ
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ Δh 283 kJ

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ Δh 1411 kJ/mole
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ Δh 1560 kJ/mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ Δh 285 kJ/mole

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 Δh is independent of the path that a
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mole of h₃o and 1 mole of oh combine
to form 2 moles of h₂o h 55 8 kj as
shown note that the enthalpy change h
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change in chemical reactions in this lab three experiments were conducted in order to test hess's law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is

to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess's law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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