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the final answer

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half reactions 6e 14h⁺ cr2o7²⁻ 2cr3⁺ 7h2o
fe2⁺ fe3⁺ e 2 equalize the electrons 6e 14h⁺ cr2o7²⁻ 2cr3⁺ 7h2o 6fe2⁺ 6fe3⁺ 6e multiplied by a factor of 6

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assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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