

Blake And Moutons Managerial Leadership Grid

Léo Pomar

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 balance the equations below
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 $3\text{KClO}_3 \rightarrow 3\text{KCl} + 2\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{Al} + 3\text{Br}_2 \rightarrow 3\text{Al}_2\text{Br}_3$
 $4\text{K} + 6\text{H}_2\text{SO}_4 \rightarrow 4\text{K}_2\text{SO}_4 + 3\text{H}_2 + 2\text{H}_2\text{O}$
 $8\text{C} + 5\text{O}_2 \rightarrow 4\text{CO}_2$

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 $1\text{O}_2 + 2\text{NH}_3 \rightarrow \text{H}_2\text{O} + 2\text{HNO}_3$
 $2\text{O}_2 + 3\text{H}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
 $4\text{RbH} + 2\text{O} \rightarrow \text{RbOH} + 2\text{H}_2\text{O}$
 $5\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$
 $6\text{N}_2 + 2\text{H}_2\text{O} \rightarrow 3\text{N}_2 + \text{H}_2$
 $3\text{Ca} + 2\text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$
 $8\text{CH}_4 + 2\text{O} \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
 5OH

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 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{Pb} + \text{OH} \rightarrow \text{H}_2 + \text{PbCl}_2$
 $3\text{Al} + 3\text{Br}_2 \rightarrow 3\text{Al}_2\text{Br}_3$
 $4\text{K} + 6\text{H}_2\text{SO}_4 \rightarrow 4\text{K}_2\text{SO}_4 + 3\text{H}_2 + 2\text{H}_2\text{O}$
 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
 $2\text{H}_2\text{O}$

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 $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O} + 2\text{C}_2\text{H}_4\text{O}_2$

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 $38\text{Fe} + 2\text{O}_2 \rightarrow 3\text{Fe}_2\text{O}_3$
 $2\text{Fe} + 3\text{CO} \rightarrow 2\text{Fe}_3\text{C} + 3\text{O}_2$
 $3\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $4\text{Mg} + \text{N}_2 \rightarrow 2\text{Mg}_3\text{N}_2$
 $4\text{P} + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$
 $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$
 $3\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $4\text{HNO}_3 \rightarrow 2\text{H}_2\text{O} + 4\text{H}_2\text{O} + \text{O}_2$

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section 4 1 is balanced meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides this is a requirement the equation must satisfy to be consistent with the law of conservation of matter

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