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Joacim Rocklöv

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example is the molar volume of a gas at stp standard
temperature and pressure which is equal to 22.4 l for 1 mole
of any ideal gas at a temperature equal to 273.15 k and a
pressure

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molar mass 32 g of the liquid expressing the density in liters
instead of ml we have $v = m \cdot d$ $\frac{32 \text{ g}}{\text{mol}} \cdot 1.790 \text{ g l}^{-1} = 0.0405 \text{ l}$
 mol^{-1} nonnumber

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volume is directly proportional to the number of moles the
important point to keep in mind is that n can be moles of any
gas since one mole always contain $6.02 \cdot 10^{23}$ molecules and
one molecule of any gas at the same temperature will have
the same kinetic energy based on this principle it is found
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one mole is defined as $6.02 \cdot 10^{23}$ particles this is the avogadro
number an unchanging constant whether you are counting
hydrogen atoms carbon dioxide molecules or carbonate ions
a mole is always $6.02 \cdot 10^{23}$ but as different particles have
different mass their molar mass varies

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though and that calculation would be exactly the one you ve
described placing x moles of a gas in a container of volume v
yields a concentration $c = \frac{x}{v}$ with units of $\frac{\text{g}}{\text{pu mol over dm}^3}$

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web the molar volume has the si unit of cubic metres per
mole $\text{m}^3 \text{ mol}^{-1}$ although it is more typical to use the units
cubic decimetres per mole $\text{dm}^3 \text{ mol}^{-1}$ for gases and cubic
centimetres per mole $\text{cm}^3 \text{ mol}^{-1}$ for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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