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A net ionic equation is the most accurate representation of the actual chemical process that occurs. Writing these equations requires a familiarity with solubility rules acid base reactivity

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Ions that remain essentially unchanged during a reaction are called spectator ions. This means you can ignore them when you write the ionic equation you only need to model how the solid

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 $3\text{H}^+ + \text{aq} + \text{AlPO}_4 + \text{s} + 3\text{K}^+ + \text{aq} + \text{net ionic} + \text{aq} + \text{PO}_4^{3-} + \text{aq} + \text{AlPO}_4 + \text{s} + 2\text{H}_2\text{O} + \text{aq}$   
 $\text{FeCl}_3 + \text{aq} + 3\text{Mg} + \text{s}$

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 the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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