

# Atatr N Genlie Hitabesi

SA Dillow

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solid is

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ternary salts consisting of ca no3 2

nano3 and kno3 the addition of

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web the amount of kno3 in the solution

is called solubility if the kno3 is less

than what the water can handle then it

is an unsaturated solution if the amount

of kno3 is equal to

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in 100g of water is 32 at 20 c and 109

at 60 c i have a doubt on 20g water is

evaporated from 200g kno3 statement i

m assuming that as

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exceptionally soluble in water your

solution is unsaturated a solution is

saturated if it contains the same

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of choice here will be the solubility

graph for potassium nitrate  $\text{KNO}_3$  which looks like this as you can see potassium nitrate has

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potassium nitrate has an orthorhombic crystal structure at room temperature which transforms to a trigonal system at 128 °C (262 °F) on cooling from 200 °C (392 °F) another trigonal phase forms between 124 °C (255 °F) and 100 °C (212 °F) sodium nitrate is isomorphous with calcite the most stable form of calcium carbonate whereas room temperature potassium nitrate is isomorphous with aragonite a slightly less stable polymorph of

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