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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol

and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the powdered acetamide occupies 1-2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affected by three key factors the first being intermolecular forces intermolecular forces affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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clean up solid and liquid spills

immediately true g conclusions after

conducting the boiling point

determination lab it can be concluded

that the unknown substance 193 is

isopropyl alcohol on account of its

observed boiling point of 88 c

conclusion after completing this**experiment the course hero** - Nov 11

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conclusion after completing this

experiment the melting point of three

different unknown compounds was

obtained using the mel temp apparatus

compound a had a melting point range

of 114 6 115 5 c thus revealing it to be

acetanilide compound b had a melting

point range of 122 6 124 5 c thus

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nov 27 2019 essay sample the purpose

of this experiment is to determine the

boiling points of various organic

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melting point can be used as the

identity and purity of an organic solid

however it is not the same as the

boiling point the boiling point is a

characteristic unique of pure liquids

the boiling point is the temperature at

which the vapor pressure of a liquid is

equal to the atmospheric pressure

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aug 11 2020 boiling points are very

sensitive to changes in applied pressure

so all boiling points should be reported

with the measured pressure a

compound s normal boiling point refers

to its boiling point at a pressure of 760

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experiment 11 vapour pressure**proctech 2ce3 lab** - Jul 07 2022

the purpose of the experiment is to

understand what vapour pressure is

and how it changes with temperature to

understand the relationship between

vapour pressure at a given temperature

and boiling temperature at a given

pressure to determine the vapour

pressure of a pure liquid at various

temperatures

exploring boiling points sharmin khatun

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exploring boiling points abstract in this

experiment we are given a set of

straight chain alkanes and a set of

compounds to compare by following the

trendline of the straight chain alkanes

it could be concluded that as the molar

mass of the straight chain alkanes

increases their boiling points increase

as well

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experiment

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laboratory partners samantha and

mackenzie goals □ correctly determine

the boiling point of an unknown liquid □

measure the refractive index of an

unknown □ use the gas chromatogram

to determine the mass percentage of a

mixture and

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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