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 KBr 2 moles times molar mass equals grams 131.108 mol \times 119
0023 g/mol 156 g KBr

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molarity of hexane C_6H_{14} 0.660 g/cm³ when it is mixed
with heptane C_7H_{12} 0.684 g/cm³ in a 35.0550 ml ratio
assume that the volumes are additive

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web notice that the mole fraction has no units on it and is
written as a decimal value do not change it to percent note of
caution you could see the term mole percent it is simply the
mole fraction multiplied by 100 for example in the problem
just below the mole fraction of cinnamic acid is 0.2885 its
mole percent would be 28.85

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percents of a mixture equals 100 mole fraction properties
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some of the other units of concentration unlike molarity mole
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