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**Ensheng Dong**

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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^{\circ}\text{C}$   $q_{\text{water}} = 550 \text{ J}$   $\Delta h_{\text{q water}} = 550 \text{ J}$

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web use the information in this chart to answer the following questions 1 if i have  $125 \text{ grams}$  of water at a temperature of  $25.0^{\circ}\text{C}$  and heat it to  $145.0^{\circ}\text{C}$  what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have  $18.7 \text{ grams}$  of an unknown liquid chemical at a temperature of  $35.0^{\circ}\text{C}$

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$\text{ml}$  of  $1.00 \text{ M}$   $\text{NaOH}$  at  $25.0^{\circ}\text{C}$  in a foam cup calorimeter a reaction occurs and the temperature rises to  $32.0^{\circ}\text{C}$  calculate the enthalpy change in  $\text{kJ}$  for this reaction assume the densities of the solutions are  $1.00 \text{ g/ml}$  and that

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from 19 77 c to 33 90 c calculate e and h for the following reaction as written 2 nh<sub>3</sub> g 3 o<sub>2</sub> g 2 h<sub>2</sub>o g

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c h 1310 kj

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