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up of copper which is put in the calorimeter are the same the
initial temperature of the calorimeter and water is 30 c and
that of hot object is 60 c the specific heat of copper and
water are 0 09 cal gm c and 1 cal gm c respectively

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temperature change plugging in the values given in the
problem you get $q_{\text{water}} = 418 \text{ J/g} \times 110 \text{ g} \times 26.6^\circ\text{C}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h_{\text{water}} = 550 \text{ J}$

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t q heat m mass qrxn qsur t tf ti name per date c specific
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is increased by 8 1 oc when 207 j of heat is added 2

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released or absorbed at constant pressure is equal to Δh the
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 $q_{\text{calorimeter}} = mcs\Delta t$ $73147314 \Delta h_{\text{rxn}} = q_{\text{rxn}}$ $q_{\text{calorimeter}} = mcs\Delta t$

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be 2 sketch the heating curve temperature vs time for the
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chemical at a temperature of 350 c

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ml of 1 00 m naoh at 25 0 c in a foam cup calorimeter a
reaction occurs and the temperature rises to 32 0 c calculate
the enthalpy change in kj for this reaction assume the
densities of the solutions are 1 00 g ml and that

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 $^\circ\text{C}$ is cooled to ice at 14 $^\circ\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^\circ\text{C}$ to steam at 138 $^\circ\text{C}$ a sample of water absorbs 1 41 $\times 10^5$ J find the mass of the sample answers 1 2 68 $\times 10^4$ J 2 6 23 $\times 10^4$ J 3 1 71 $\times 10^4$ J 4 1 11 $\times 10^5$ J 5 2 11 $\times 10^5$ J 6

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