

# Fruits Basket Volume 23

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web 1 if the pe of the reactants is 430kj and the pe of the products is 190kj what is the  $\Delta h$  2 if the pe of the reactants is 430kj and the activation energy is 220 how much total energy will be needed to be added for the reaction to

proceed transcribed image text potential energy diagram  
answer the questions using the graph

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web the energy of the reactants is shown on the energy diagram above on the right side of the energy diagram draw a horizontal line segment to indicate the energy of the products draw a vertical line segment to indicate  $\Delta h$  for the reaction and label it with the correct value see attached image

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web aug 16 2019 1 answer sorted by 4 the activation energy is not the average kinetic energy that the colliding particles need in order to successfully react instead it is the energy barrier between reactants and products often the energy required to cross the barrier comes from the kinetic



energy of collisions

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diagrams for exothermic and endothermic reactions and

what effect adding a catalyst or inhibitor has on the diagram

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web 1 potential energy diagrams are graphs energy is on the

y axis and the x axis indicates geometry energy is a function

of geometry 2 the geometry changes refer to changes in

bonding changes in atom positions each point on the diagram

has the same molecular formula same atoms and electrons

the bonding can change however

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reaction recall that the enthalpy change  $\Delta h$  is positive for an

endothermic reaction and negative for an

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reaction 2 label your drawing with the following letters a pe

of products d energy of activation b pe of reactants e heat of

reaction c pe of the activated complex potential energy

diagrams time ms 0 00 1 00 2 00 3 00 4 00 5 00 6 00 7 00 8

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