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Lauren Gardner

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moles of butane in order to find the molarity the mass of the
butane should be divided by the moles after further
calculations the molarity of the butane gas was calculated to
be 56 81 grams moles this lab is done with tons of sources of
error

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web this data shows that there was a gain the mass of the
lighter which shows that there was a certain amount of

butane 0 38 grams 0 02 grams gas that was successfully
collected during the gas collection there is a mild influence
of barometric pressure during both trials as shown above in
the raw data

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of butane gas assume that butane is an ideal gas and that
one mole has a volume of 22 4 l at stp v 1 n 1 v 2 n 2 0 0896 l
n 22 4 l 1 mole n 0 0040 moles 12 experimental molar mass g
mole of butane 0 23 g 0 0040 moles 58 g mole 13 the
molecular formula of butane is c 4 h 10

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mass of a butane lighter was measured before using the
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