

Castro De Neixon Boiro

N Colangelo

How Do We Imagine the Past? On Metaphorical Thought, Experientiality and Imagination in Archaeology Paul Bouissac, Dragoş Gheorghiu, 2015-02-27 Recent years have witnessed a search for new sources for archaeological inspiration within areas which until recently have not been imagined as a source for science. Archaeology has become more “anthropologized”, and, as such, is becoming increasingly influenced by the Zeitgeist, although some European schools are yet to recognize this. The process of scientific research that archaeologists have always considered to be an objective approach has been revealed to be the result of different subjective cognitive processes, forming part of the contemporary humanistic paradigm, a fact confirmed by new tendencies in contemporary archaeology. Consequently, this book considers the question: how does the archaeologist think today? Beginning with simple analogies issued from archaeological experiments or from ethnography, the structure of the contemporary archaeological thought is increasingly complex, working today with concepts that only yesterday were a subject of study. This book considers these new types of approaches, through a series of personal narratives provided by archaeologists, describing their working methods in the process of imagining the past.

Cuadernos de estudios gallegos ,2008

Fasti Archaeologici ,1991 Annual bulletin of classical archaeology.

Humanitas Antón A. Rodríguez Casal,1996

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Arqueología 80 España. Subdirección General de Arqueología y Etnografía,1980

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them

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$1\text{K}_3\text{PO}_4 + \text{Al}(\text{NO}_3)_3 + 3\text{KOH} \rightarrow 3\text{K}^+ + 3\text{PO}_4^{3-} + \text{Al}^{3+} + 3\text{OH}^-$ net ionic $\text{Al}^{3+} + 3\text{OH}^- \rightarrow \text{Al}(\text{OH})_3$

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