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 where is the potential energy where is the kinetic energy  
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 the reaction kinetic energy is the heat that is either released  
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positive negative  $h$  is for endothermic reactions and for exothermic reactions enthalpy is a state function and is therefore independent of path state hess's law enthalpy of formation  $h_f$  the  $h$  of a formation reaction  $h_f$  products  $h_f$  reactants

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