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solution $4\text{CN}^- + 2\text{Au} + \text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{H}_2\text{O} + \text{O}_2 + 2\text{H}^+ + 3\text{e}^-$ the final answer

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ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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cr2o72 fe2 cr3 fe3 solution 1 balanced

half reactions 6e 14h cr2o72 2cr3 7h2o fe2 fe3 e 2 equalize the electrons 6e 14h cr2o72 2cr3 7h2o 6fe2 6fe3 6e multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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