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 $2 \text{ICl(g)} + \text{H}_2\text{(g)} \rightarrow 2 \text{HCl(g)} + 2 \text{I(g)}$
 a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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 reaction below if substance a is
 disappearing at a rate of 1.82×10^{-2} mol l⁻¹ s⁻¹ at what rate is c appearing
 3 a 3 b \rightarrow 5 c 2 d
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 o3 3 o2 b identify any intermediates
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 mg s 2 hcl a q mgcl a 2 a q h a 2 g the
 reaction between mg s and hcl a q is
 represented by the equation above in a
 kinetics experiment a 0.080 g sample of
 mg s ribbon is added to an erlenmeyer
 flask containing 50 ml of 3.0 m hcl a q
 at 25 c

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