

Ppt On Effective Teaching Strategies

Marcel A. Müller

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experimental light brown nutmeg powder 4 g was mixed with 25 ml of clear and colorless diethyl ether solvent and the brown cloudy mixture was filtered by gravity filtration with fluted fast flow filter paper in a powder funnel into a 250 ml round bottom flask rbf leaving light brown residue and beige filtrate diethyl ether was

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 of the organic chem lab survival manual
 prelab question 1 trimyristin is a
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 approximately 2.5 g of ground nutmeg
 record exact mass and 15 ml of diethyl
 ether to microwave vessel
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 purpose of this lab is to new techniques
 isolate the chemical trimyristin from the
 common spice nutmeg this lab requires
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 over 40 of the medicinal chemicals used
 throughout the developed world today
 were originally isolated from natural
 sources these sources include flowering
 plants fungi bacteria and to a lesser
 extent animals especially marine
 animals
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 trimyristin will be useful for students to
 obtain natural whitening raw materials
 derived from nutmeg various studies on
 the isolation can be used in natural
 products
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lot easier when extracting trimyristin
 from nutmeg because of its large
 quantity in nutmeg the nutmeg is
 moved from a solid to a liquid phase
 which is then used for extracting and
 isolating the trimyristin
development of laboratory module of
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 web trimyristin compound from nutmeg
 seed trimyristin compounds contained
 in nutmeg seed is very important to be
 isolated because it has various benefits
 trimyristin along with myristic acid
 myristicin and elemicin have
 antioxidant anticonvulsant analgesic
 anti inflammatory antidiabetic
 antibacterial and anti fungal activity
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 web aug 4 2008 the purpose of this
 experiment was extract isolate and
 purify the natural product trimyristin
 from the spice nutmeg the trimyristin
 was extracted by using the solvent
 diethyl ether and
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of nutmeg the fixed oil comprises approximately 24.40% of the nutmeg seed. Trimyristin comprises 73% of the fixed oil. Overall, trimyristin should have a percent recovery of 18.291%. Figure 1 shows how trimyristin is triester formed from the dehydration reaction between glycerol and myristic acid.

Isolation of trimyristin from nutmeg
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 web purpose the purpose of this lab is to extract trimyristin from nutmeg as well as synthesizing myristic acid from the extracted trimyristin using base hydrolysis and NaOH acidification. Using HCl is necessary to characterize the product by finding the melting point.

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 web in this lab trimyristin is extracted from ground nutmeg occurring in many vegetable oils as well as myristic acid. Trimyristin consists of long saturated hydrocarbon chains and it is relatively nonpolar. Thus, using a relatively nonpolar solvent, trimyristin can be easily extracted by the concept of like dissolves like.

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 web the desired components of solids are dissolved in a solvent. Then the remaining solid is filtered so that the selected element and solvent are left over. In this lab, the trimyristin was dissolved into the solvent and the remainder of the nutmeg was filtered out.

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