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 2 requires 2 m of n a oh per liter so 4 x
 2 8 m 4 a a solution of molarity 1 5 m
 requires 1 5 mol of na to every litre of
 solvent

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 $\frac{1.5 \times 480 + 1.2 \times 520}{480 + 520}$
 $\frac{720 + 624}{1000} = 1.344$ m q5 with increase
 in temperature which of these changes
 a molality b fraction of solute present in
 water

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web 1 calculate the molarity of a
 solution which contains 0 40 mol of a
 substance dissolved in 1 6 l of a
 solution 2 what is the molarity of a
 solution containing 325 g of nacl
 dissolved in 750 ml of solution 3 140 g
 of kcl is dissolved in 600 ml of water
 what is the molarity 4 724 4 g of
 ammonium phosphate in 4500 ml of
 alcohol

6 1 calculating molarity problems

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web determine the molarity for each of
 the following solutions 0 444 mol of
 cocl₂ in 0 654 l of solution 98 0 g of
 phosphoric acid h₃ po₄ in 1 00 l of
 solution 0 2074 g of calcium hydroxide
 ca oh₂ in 40 00 ml of solution 10 5 kg

16 11 molality chemistry libretexts -

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web molality differs from molarity only
 in the denominator while molarity is
 based on the liters of solution molality
 is based on the kilograms of solvent
 concentrations expressed in molality
 are used when studying properties of
 solutions related to vapor pressure and
 temperature changes

molarity and molality practice

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web molarity molality notes and
 practice answer the questions below
 show all work including units watch
 your significant digits and circle your
 answers molarity just a reminder
 molarity is one of the many ways to
 measure concentration or