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phosphoric acid h₃po₄ in 1 00 l of

solution 0 2074 g of calcium hydroxide

ca oh₂ in 40 00 ml of solution

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c₆h₁₄ d 0 660 g cm³ when it is

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cm³ in a 35 0 55 0 ml ratio assume that

the volumes are additive

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to use the equation that says that the

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equation to find that the concentration
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containing 571 4 g of h₂so₄ per liter

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molarity of each of the following

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in 0 100 l of serum the average

concentration of cholesterol in human

serum b 4 25 g of nh₃ in 0 500 l of

solution the concentration of nh₃ in

household ammonia

15 03 solution concentration

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web a similar unit of concentration is

molality m which is defined as the number of moles of solute per kilogram of solvent not per liter of solution
 molality $\frac{\text{moles solute}}{\text{kilograms solvent}}$
 mathematical manipulation of molality is the

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web 1 molarity m $\frac{\text{mass molar mass} \times 0.4500}{10.825 \text{ g} + 141.9579 \text{ g mol} \times 0.0129}$
 m 2 molality $\frac{0.825 \text{ g} + 141.9579 \text{ g mol} \times 0.00581158}{0.00581158 \text{ mol} + 0.4500 \text{ kg} + 0.0129 \text{ m} + 3 \text{ mole fraction na} + 2 \text{ hpo} + 4.0825 \text{ g} + 141.9579 \text{ g mol} + 0.00581158 \text{ mol} + 2.0450 \text{ g} + 18.015 \text{ g mol} + 24.97918401 \text{ mol}}$

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web molality differs from molarity only in the denominator while molarity is

based on the liters of solution molality is based on the kilograms of solvent
 concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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concentration of Ca^{2+} ions in a water sample with a hardness count of 175 $\text{mg CaCO}_3 \text{ l}^{-1}$

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it is written like this
 begin array $\frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution} \times 100}$ end array

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