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web donor whereas a base generically rep resented by b is a proton acceptor when acids and bases combine they swap a proton ha b a hb strong acids e g sulfuric acid have a very pronounced

ten dency to transfer a proton to a base
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web understand that parts of an acid

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k a or k b from ph data associated with

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able to determine the molar mass of a

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data be able to calculate k a1 and k a2

for a polyprotic acid by the end of this

lab students should be able to

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the titration the solution in the

erlenmeyer flask is acidic as the base is

added it completely reacts with the acid

and the solution in the erlenmeyer flask

continues to be acidic but at the

equivalence point

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strong acid strong base combination of

hydrochloric acid and sodium hydroxide

had the neutral ph at the equivalence

point being 7 this result is expected

because a combination of strong acids

and strong bases should result in a

neutral at the equivalence point

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laboratories to determine the quantity

or the concentration of a known or

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technique used to determine the

volume of a solution that is needed to

react with a given amount of another

substance in this activity your goal is to

determine the molar concentration and

strength of two acid solutions by

conducting titrations with a base

solution of known

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oxalic acid h 2 c 2 o 4 2h 2 o was

weighted to the nearest 0 g in a 50 ml

beaker 2 about 30 ml of distilled water

was added to the beaker to dissolve the

oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = \log 0 100 1 000$ ph $\log 0 100 1 000$ b titrant volume 12 50 ml

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