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 q as negative when heat flows from the
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 Δh of reaction the word exothermic is
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 from the acid and the oh^- ions from the
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neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to

produce one mole of water for example
 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$ you can

see that one mole of water is produced when one mole of hcl or one mole of naoh reacts