

I Gemelli Nati Dalluovo Di Leda

A Loxley

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subsequent ΔH values
 $2 \text{ CO}_2 \text{ g} \rightarrow \text{H}_2\text{O g}$
 $\text{C}_2\text{H}_2 \text{ g} \rightarrow \text{H}_2\text{O g}$
 $\text{C}_2\text{H}_2 \text{ g} \rightarrow \text{H}_2 \text{ g}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \Delta H$ 94
 5 kJ
 $\text{H}_2\text{O g} \rightarrow \text{H}_2 \text{ g}$
 $\frac{1}{2}\text{O}_2 \text{ g} \rightarrow \Delta H$ 71
 2 kJ
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \frac{7}{2}\text{O}_2 \text{ g}$
 $2 \text{ CO}_2 \text{ g} \rightarrow 3 \text{ H}_2 \text{ O g}$
 ΔH 283 kJ
 g
 $\text{C}_2\text{H}_6 \text{ g}$

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 worksheet calculate ΔH for the reaction
 $\text{C}_2\text{H}_4 \text{ g} \rightarrow \text{H}_2 \text{ g}$
 $\text{C}_2\text{H}_6 \text{ g}$ from the
 following data
 $\text{C}_2\text{H}_4 \text{ g} \rightarrow 3 \text{ O}_2 \text{ g}$
 $2 \text{ CO}_2 \text{ g} \rightarrow \text{H}_2 \text{ O l}$ 1411 kJ
 mole
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow 7 \text{ O}_2 \text{ g}$
 $2 \text{ CO}_2 \text{ g} \rightarrow 3 \text{ H}_2 \text{ O l}$ 1560 kJ
 mole
 $\text{H}_2 \text{ O l} \rightarrow \text{H}_2 \text{ O g}$ 285 kJ
 mole

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 hess s law example 12 1 determine ΔH for the target reaction
 $\text{C}_2\text{H}_4 \text{ g} \rightarrow \text{H}_2\text{O g}$
 $\text{C}_2\text{H}_6 \text{ g}$ given the following information

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 libretexts *quot hess s law quot* 2019 which determines the enthalpy change value
 hess s law is written as
 $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$ or
 $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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$\text{Mg s} \rightarrow \frac{1}{2}\text{O}_2 \text{ g}$
 MgO s we used
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 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH s} \rightarrow \text{Na}^+ \text{aq}$
 $\text{OH}^- \text{aq}$

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 enthalpy change ΔH is independent of the path that a reaction follows to move from reactants

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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temperature and the c given trial 1 use hess s law and the following equations and ΔH values to determine the heat of reaction for the reaction C_2H_4

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$ $\text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ $\text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$

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admonish advise caution sentinel
person or animal that guards mead
liquor made from fermented honey
water fervent having or showing great
warmth of feeling rancor ill will

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mythology accumulation of wisdom of a

given society themes involved birth life
and death great search mythical
wandering of an individual on a quest

which involves monsters enchantresses
and an experience with the world of the
dead often the hero is trying to