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web donor whereas a base generically represented by  $b$  is a proton acceptor when acids and bases combine they swap a proton  $h_a$   $b$  a  $h_b$  strong acids e.g sulfuric acid have a very pronounced

tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point [pdf acid base titration researchgate](#) - Nov 27 2022

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighed to the nearest 0 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the

oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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000  $\text{ph} = \log 0 100 1 000$  b titrant volume 12 50 ml

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